

Name _____ Period _____ Date _____

World of Chemistry

Watch both videos and answer questions in homework section of lab notebook

Episode 13 - *The Driving Forces*

1. Describe the energy levels of reactants and products in an exothermic reaction.
2. In general, most chemical reactions _____ energy.
3. Why does an endothermic reaction feel cold if energy is being absorbed?
4. Describe the energy levels of reactants and products in an endothermic reaction.
5. Give two everyday examples of endothermic reactions.
6. What is meant by the term entropy?
7. In which of the following examples is entropy INCREASING?
 - a. solid turns to liquid
 - b. liquid turns to gas
 - c. solid turns to gas
8. In general, reactions proceed toward _____ energy and _____ entropy.
9. What three factors were mentioned that determine how fast a chemical reaction takes place

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Episode 14 - *Molecules in Action*

1. What is the basis of all chemical reactions?
2. What is necessary for successful collisions?
3. What is meant by activation energy?
4. How are activation energies affected by catalysts?
5. What is true about a catalyst at the end of a chemical reaction?
6. What is meant by equilibrium?
7. What does Le Chatelier's Principle tell us about equilibrium reactions?
8. What is the reaction involved in the Haber Process?
9. Why is this reaction important?
10. What conditions are used to increase the yield in the Haber Process?

